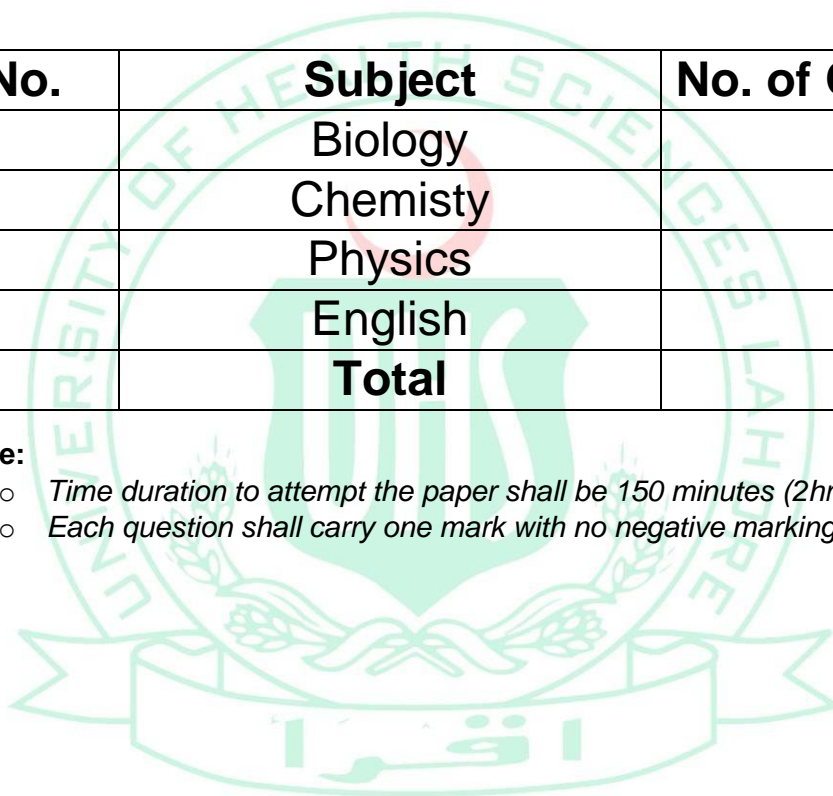


**Structure of the Question Paper for  
Medical and Dental Colleges Admission Test  
(MDCAT) 2020**

<b>Sr. No.</b>	<b>Subject</b>	<b>No. of Questions</b>
1	Biology	80
2	Chemistry	60
3	Physics	40
4	English	20
	<b>Total</b>	<b>200</b>

- **Note:**
  - Time duration to attempt the paper shall be 150 minutes (2hrs 30 minutes).
  - Each question shall carry one mark with no negative marking.



- 
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# BIOLOGY

## Table of Contents

- 1 The Cell
- 2 Biological Molecules
- 3 Chromosomes and DNA
- 4 Cell Division
- 5 Variety of Life
- 6 Bioenergetics
- 7 Gas Exchange
- 8 Transport in Plants
- 9 Transport in Human
- 10 Immunity
- 11 Homeostasis
- 12 Muscles and Movement
- 13 Communication
- 14 Reproduction
- 15 Genetics
- 16 Biotechnology
- 17 Evolution



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# 1 The Cell

## COURSE CONTENT

- Light and Electron Microscope (Magnification and Resolution)
- Structure of Typical Animal and Plant Cell
- Fluid Mosaic Model of Cell Membrane
- Transport of Material across the Cell Membrane: Active transport, Passive transport, Endocytosis and Exocytosis
- Eukaryotic Cell Structures: Endoplasmic reticulum (RER & SER), Ribosomes, Golgi apparatus, Lysosomes, Vacuoles, Centrioles & Microtubules, Mitochondria, Chloroplast and Nucleus (nuclear membrane, nucleolus and chromosomes)
- Prokaryotic Cell & Eukaryotic cell

# 2 Biological Molecules

## COURSE CONTENT

- Carbohydrates: Monosaccharides, Disaccharides and Polysaccharides (Starch, Glycogen & Cellulose)
- Lipids: Triglycerides, Phospholipids and their functions.
- Proteins: Amino Acids & Peptide bond formation, Structures of Proteins (primary, secondary, tertiary and quaternary structures) and Globular & Fibrous Proteins
- Nucleic acids: DNA, RNA and Types of RNA
- Water: Heat of vaporization, Specific Heat Capacity and Solvent Action
- Enzyme: Definition, Characteristics of, Mechanism of Enzyme Action Enzymes (Lock & key model and Induced fit model), Factors affecting the rate of Enzyme Action, Inhibitors

# 3 Chromosomes and DNA

## COURSE CONTENT

- Chromosome: Nucleosome, DNA, Histone Proteins, Chromatids, Centromere and Telomeres

- 
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- Gene as a Basic Unit of Genetic Information
- DNA Replication: Hypothesis of DNA Replication, Meselson & Stahl's experiment and Replication
- Transcription
- Genetic Code
- Translation

## 4 Cell Division

### COURSE CONTENT

- Cell Cycle: Interphase (G1, S and G2 phases), Mitotic phase and Cytokinesis
- Mitosis: Process of Mitosis, Significance of Mitosis
- Meiosis: Process of Meiosis and Significance of Meiosis

## 5 Variety of Life

### COURSE CONTENT

- Kingdoms: Protoctista, Fungi, Plantae and Animalia
- Viruses: Structure of Viruses
- AIDS: Causative Agent, Modes of Transmission and Prevention & Control

## 6 Bioenergetics

### COURSE CONTENT

- Photosynthetic Pigments (Chlorophylls and Carotenoids)
- Absorption and Action Spectra
- Light-Dependent Reactions (cyclic and non-cyclic phosphorylation) and Light-Independent Reactions (Calvin cycle).
- Cellular Respiration: Glycolysis, Link reaction / Pyruvic Oxidation, Krebs's Cycle (with reference to production of NADH, FADH and ATP) and ETC
- Anaerobic Respiration and its Types (Alcoholic and Lactic Acid Fermentation).

- 
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## 7 Gas Exchange

### COURSE CONTENT

- Anatomy of Human Respiratory System.
- Transport of Respiratory Gases: O<sub>2</sub> & CO<sub>2</sub> and Role of Haemoglobin as Respiratory Pigment.
- Respiratory Disorders: Tuberculosis, Emphysema and Lung Cancer

## 8 Transport in Plants

### COURSE CONTENT

- Transport of Water and Minerals: Apoplast & Symplast Pathway and Cohesion, Transpiration Pull / Tension & Adhesion
- Transpiration, Factors affecting it and opening and closing of Stomata.
- Translocation according to Pressure Flow Theory
- Xerophytes

## 9 Transport in Human

### COURSE CONTENT

- Heart: Structure of heart, Cardiac Cycle, Control of Heart Beat, ECG and Blood Pressure
- Blood Vessels: Arteries, Veins and Capillaries
- Blood: Plasma and Blood Cells (RBCs, WBCs and platelets).
- Lymphatic System

## 10 Immunity

### COURSE CONTENT

- Immune System and its Components
- Types of Immunity
- Vaccination

- 
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## 11 Homeostasis

### COURSE CONTENT

- Homeostasis
- Thermoregulation in Mammals
- Human Urinary System

## 12 Muscles and Movement

### COURSE CONTENT

- Structure and Function of Skeletal Muscle
- Mechanism of Skeletal Muscle Contraction; Sarcomere, Ultrastructure of Myofilaments, Sliding Filament, Control of Actin-Myosin Interaction and Use of Energy for Muscle Contraction.

## 13 Communication

### COURSE CONTENT

- Nervous Coordination in Mammals
- Neurons: Sensory, Intermediate / relay and motor neurons
- Reflex arc / Reflex action
- Nerve impulse
- Synapse
- Hormones: Definition & Types of Hormones, Hormones of Islets of Langerhans (Insulin & Glucagon) and Role of ADH in Osmoregulation.
- Plants Hormones: Auxins, Gibberellins and Abscisic Acid

## 14 Reproduction

### COURSE CONTENT

- Gametogenesis: Spermatogenesis and Oogenesis
- Hormonal control of Human Menstrual Cycle (FSH, LH, estrogen and progesterone)

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## 15 Genetics

### COURSE CONTENT

- Basics of Genetics: Gene, Locus, Allele, Gene Pool, Phenotype, Genotype, Homozygous, Heterozygous, Dominant Allele, Recessive Allele, Complete Dominance, Codominance, Linkage, F<sub>1</sub> & F<sub>2</sub> Generations, Mutation and Multiple Allele.
- Gene Linkage: Crossing over and Recombination Frequency / Cross Over Value
- Continuous and Discontinuous Variations
- Punnet square, Test cross and Monohybrid & Dihybrid Crosses
- Gene Linkage and Sex Linkage in Human (Haemophilia and Colour Blindness)

## 16 Biotechnology

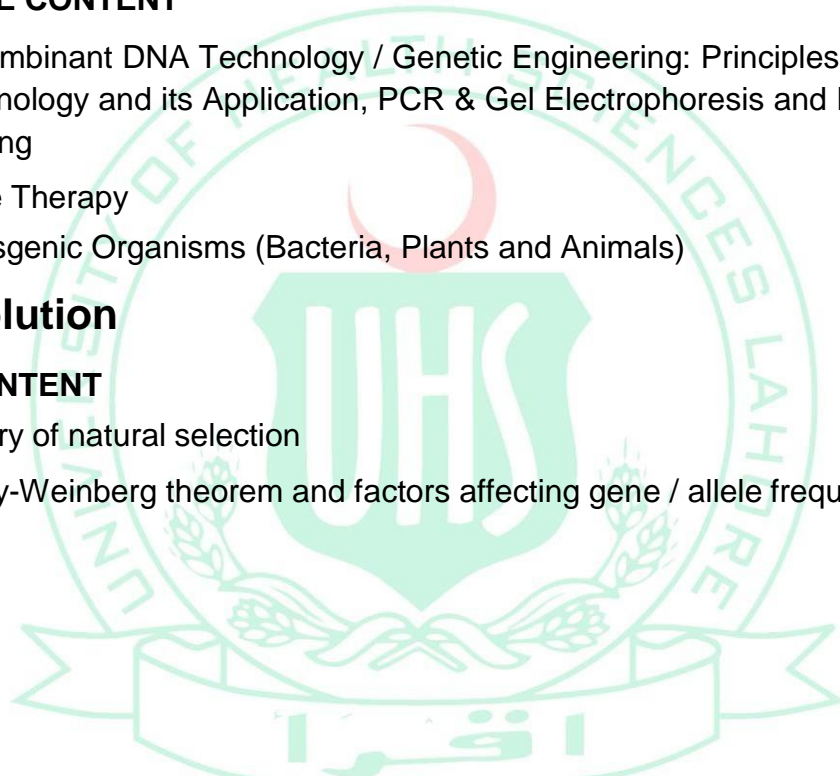
### COURSE CONTENT

- Recombinant DNA Technology / Genetic Engineering: Principles of Recombinant DNA Technology and its Application, PCR & Gel Electrophoresis and DNA Analysis / Finger Printing
- Gene Therapy
- Transgenic Organisms (Bacteria, Plants and Animals)

## 17 Evolution

### COURSE CONTENT

- Theory of natural selection
- Hardy-Weinberg theorem and factors affecting gene / allele frequency



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# CHEMISTRY

## TABLE OF CONTENTS

### A. Physical Chemistry

1. Fundamental Concepts
2. States of Matter
3. Atomic Structure
4. Chemical Bonding
5. Chemical Energetics
6. Electrochemistry
7. Chemical Equilibrium
8. Reaction Kinetics

### B. Inorganic Chemistry

1. Periods
2. Groups
3. Transition Elements
4. Compounds of Nitrogen and Sulphur

### C. Organic Chemistry

1. Fundamental Principles
2. Hydrocarbons
3. Alkyl Halides (Haloalkanes)
4. Alcohols and Phenols
5. Aldehydes and Ketones
6. Carboxylic Acids
7. Amino Acids
8. Macromolecules
9. Environmental Chemistry

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# PHYSICAL CHEMISTRY

## 1. FUNDAMENTAL CONCEPTS

### COURSE CONTENT

- Relative masses
- Isotopes
- Mole
- Avogadro's number
- Empirical and Molecular formulae
- Stoichiometric Calculations
- Concentration units of solutions
  - i. Percentage composition
  - ii. Molarity
  - iii. Mole fraction

## 2. STATES OF MATTER

### COURSE CONTENT

- Gaseous state
- General Gas Equation ( $PV=n RT$ ).
- Liquid state:
  - i. Evaporation
  - ii. Vapor pressure
  - iii. Boiling
  - iv. Structure of ice
- Lattice structure of a Crystalline solid

## 3. ATOMIC STRUCTURE

### COURSE CONTENT

- Proton, Neutron and Electron
- Distribution of Mass and charges
- Deduce the number of protons, neutrons and electrons from given proton number and nucleon number

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- Shape of s, p and d- Orbitals
- Electronic configuration
- Ionization energy
- Electron affinity

#### 4. CHEMICAL BONDING

##### COURSE CONTENT

- Ionic (Electrovalent) bond
- Use the “dot and cross” model for:
  - i. Covalent bonding
  - ii. Co-ordinate (dative covalent) bonding
- Shapes and Bond Angles of molecules
- Covalent Bonding.
- Bond Energy, Bond length and bond Polarity (Electronegativity difference)
- Intermolecular Forces (especially Hydrogen Bonding)
- Interpret and Predict the effect of different types of bonding on physical properties of substances.

#### 5. CHEMICAL ENERGETICS

##### COURSE CONTENT

- Concept of Energy changes during Chemical reactions
- Use the terms:
  - i. Enthalpy change of reaction
  - ii. Bond energy
- Numerical Magnitude of Lattice Energy
- Hess’s law to construct simple energy cycles.

#### 6. ELECTROCHEMISTRY

##### COURSE CONTENT

- Redox processes
- Oxidation numbers of Elements.
- Balancing chemical equations by redox method

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- Standard electrode (redox) Potential
- Standard Hydrogen Electrode
- Methods used to measure the standard Electrode potentials of metals
- Standard Cell Potential
- Electrode Potential.
- Construct Redox equations
- Advantages of Developing the H<sub>2</sub>/O<sub>2</sub> fuel cell

## 7. CHEMICAL EQUILIBRIUM

### COURSE CONTENT

- Rates of forward and reverse Reactions and Dynamic Equilibrium
- Le- Chatelier's Principle.
- Deduce expression for Equilibrium constant
- Calculate the values of Equilibrium constants
- Calculate the equilibrium quantities
- Conditions used in Haber process.
- Qualitatively the differences in behavior of strong/weak acids and bases
- Terms: pH, K<sub>a</sub>, pK<sub>a</sub>, K<sub>b</sub>, pK<sub>b</sub>, K<sub>w</sub> and K<sub>sp</sub>.
- [H<sup>+</sup> (aq)], [OH<sup>-</sup> (aq)], pH and pOH values for strong and weak acids and bases.

## 8. REACTION KINETICS/ CHEMICAL KINETICS

### COURSE CONTENT

- Rate of Reaction, Activation Energy, Catalysis, Rate Equation, Order of Reaction, Rate Constant
- Collisions
- Enzymes as Biological Catalysts
- Construct and use rate equations with special emphasis on
  - Zero order reaction
  - 1<sup>st</sup> order reaction
  - 2<sup>nd</sup> order reaction
- Half-life of a first order Reaction

- 
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- Calculate the order of reaction
- Calculate the rate constant
- Name a Suitable method for studying the rate of a Reaction

## INORGANIC CHEMISTRY

### 1. PERIODS

#### COURSE CONTENT

- Variation in the Physical properties of Elements Belonging to period 2 and period 3:
  - a. Atomic Radius
  - b. Ionic Radius
  - c. Melting Point
  - d. Boiling Point
  - e. Ionization Energy
  - f. Electronegativity
  - g. Electron Affinity
  - h. Electrical Conductivity
  - i. Oxidation States
  - j. Hydration Energy
- Periodic Relationship in Binary compounds:
  - a. Halides (especially chlorides)
  - b. Oxides

### 2. GROUPS:

#### COURSE CONTENT

- The variation in the properties of group II and VII Elements:
  - a. Reactions of group II elements.
  - b. Thermal decomposition and solubility of compounds of group II elements
  - c. Properties of Halogens and their compounds
  - d. Reaction of Chlorine
  - e. Comparison of Oxidizing power of halogens

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- f. Uses of Halogens and compounds of Halogens

### 3. TRANSITION ELEMENTS:

#### COURSE CONTENT

- Chemistry of Transition Elements of 3d series:
  - a. Electronic Configuration
  - b. Variable Oxidation states
  - c. Uses as a Catalyst
  - d. Formation of Complexes
  - e. Colour of Transition Metal Complexes
  - f. Geometry and isomerism of complex ions with coordination number 4 and 6.

### 4. COMPOUNDS OF NITROGEN AND SULPHUR

#### COURSE CONTENT

- Nitrogenous fertilizers.
- Presence of Sulphur Dioxide in atmosphere
- Manufacturing of Sulphuric Acid.
- Sulphuric acid as Dehydrating agent and Oxidizing agent.

## ORGANIC CHEMISTRY

### 1. FUNDAMENTAL PRINCIPLES:

#### COURSE CONTENT

- The Organic compounds.
- Alkanes and Alkenes of lower masses.
- Nucleophiles, electrophiles and free radicals
- Isomerism.
- Functional group and Nomenclature of organic compounds.

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## 2. HYDROCARBON:

### COURSE CONTENT

#### Chemistry of Alkanes with emphasis on:

- Combustion.
- The Mechanism of free radical Substitution reaction.

#### Chemistry of Alkenes with emphasis on:

- Preparation of Alkenes:
- Reaction of Alkenes:

#### Chemistry of Benzene

- Benzene.
- Electrophilic substitution reactions and the mechanism
- Hydrogenation of Benzene ring.
- Side chain Oxidation of Methyl Benzene (Toluene) and Ethyl Benzene.
- Benzene ring by 2,4 Directing and 3,5 directing groups

## 3. ALKYL HALIDES (HALOGENOALKANES)

### COURSE CONTENT

- Importance of Halogenoalkanes.
- Reaction of Alkyl Halides:
- Nucleophilic substitution reactions and mechanisms

## 4. ALCOHOLS AND PHENOLS:

### COURSE CONTENT

#### Alcohols with reference to:

- Alcohols: Primary, Secondary and Tertiary.
- Preparation of Ethanol
- Reactions of Alcohols

#### Phenols

- Reactions of Phenol
- Acidity of Water, ethanol and phenol

- 
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**5. ALDEHYDES AND KETONES:****COURSE CONTENT**

- Structure of Aldehydes and Ketones.
- Preparation of Aldehydes and Ketones.
- Reactions of Aldehydes and Ketones:
- Nucleophilic addition reaction mechanism

**6. CARBOXYLIC ACID:****COURSE CONTENT**

- Physical properties of carboxylic acids.
- Reactions of carboxylic acids
- Relative Acidic strength.

**7. AMINO ACIDS:****COURSE CONTENT**

- General Structure of  $\alpha$ -Amino Acids found in Proteins.
- Amino Acids on the basis of Nature of R-group.
- Zwitter Ion
- Acid base properties of Amino Acids.
- Peptide bond formation.

**8. MACROMOLECULES:****COURSE CONTENT**

- Addition polymerization.
- Condensation polymerization.
- Structure of Proteins.
- Structure and function of Nucleic acid (DNA).

- 
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## 9. ENVIRONMENTAL CHEMISTRY:

### COURSE CONTENT

- Air Pollutants.
- Chemistry and causes of Acid Rain.
- Ozone and Chlorofluorocarbons (CFCs).



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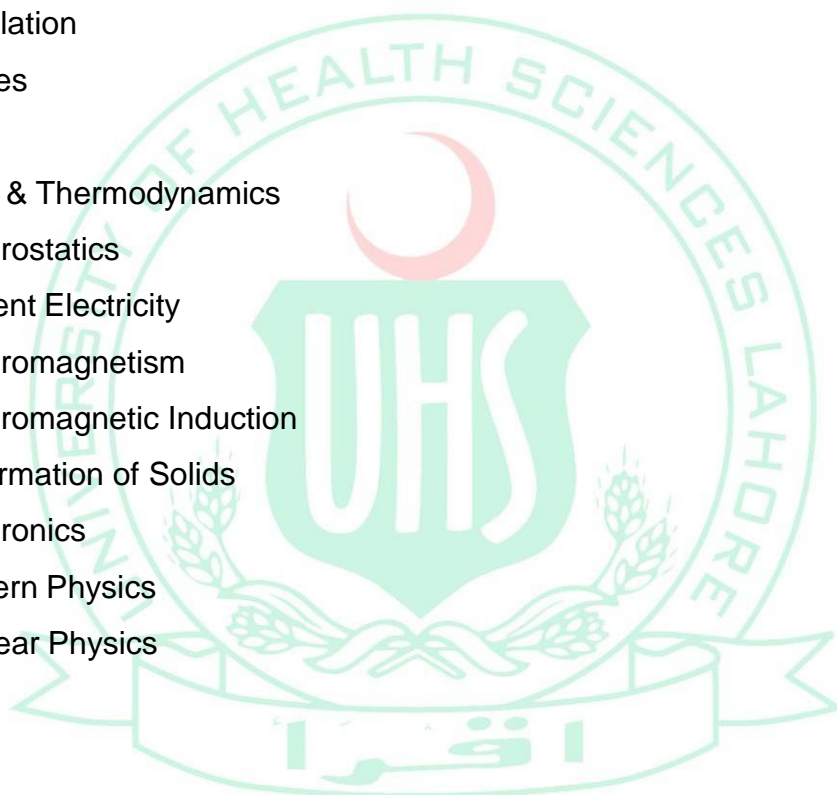
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# PHYSICS

## Table of Contents

1. Measurement
2. Motion and Force
3. Work, Energy and Power
4. Circular Motion
5. Oscillation
6. Waves
7. Light
8. Heat & Thermodynamics
9. Electrostatics
10. Current Electricity
11. Electromagnetism
12. Electromagnetic Induction
13. Deformation of Solids
14. Electronics
15. Modern Physics
16. Nuclear Physics



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## 1. MEASUREMENT

### COURSE CONTENT

- Physical quantities, numerical magnitude and a proper unit.
- International System of Units, SI base units of physical quantities, and their derived units.
- Prefixes and symbols to indicate decimal, submultiples or multiples of both base and derived units:
- Errors and uncertainties
  - Systematic error and random error.
  - Fractional uncertainty and percentage uncertainty.
  - Assessment of total uncertainty in the final results (Understanding of total assessment about addition and subtraction, multiplication and division & power factor).

## 2. MOTION AND FORCE

### COURSE CONTENT

- Displacement, Distance, Speed, Velocity and Acceleration.
- Velocity–Time Graph.
- Equations of motion.
- Newton’s Laws of Motion.
- Momentum and law of conservation of momentum.
- Force and rate of change of momentum.
- Impulse and  $I = F \times t = mv_f - mv_i$
- Elastic and in-elastic collisions
- Projectile Motion and its applications.
- Moment of force or torque and use of torque
- Equilibrium.

## 3. WORK, ENERGY AND POWER

### COURSE CONTENT

- Work in terms of the product of a force and displacement in the direction of the force.
- Kinetic energy  $K.E \equiv \frac{1}{2} mv^2$
- Potential energy  $P.E = mgh$
- Inter-conversion of kinetic energy and potential energy in gravitational field.
- Power in terms of work done per unit time and use power as product of force and

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$$\text{velocity } P = \frac{W}{t} \text{ and } P = Fv$$

## 4. CIRCULAR MOTION

### COURSE CONTENT

- Angular motion, angular displacement & angular velocity
- Centripetal Force and centripetal acceleration

$$\circ F = mr^2\omega, F = \frac{mv^2}{r} \text{ and } a_c = r\omega^2 \text{ and } a_c = \frac{v^2}{r}$$

- Geostationary orbits.
- Radian

## 5. OSCILLATIONS

### COURSE CONTENT

- Simple harmonic motion
- Amplitude, Frequency, Angular Frequency, Phase Difference. Express the time period in terms of both frequency and angular frequency.
- Equations  $x = x_0 \sin \omega t$ ,  $v = v_0 \cos \omega t$ ,  $v = \sqrt{x_0^2 - x^2}$ ,  $a = -\omega^2 x$  and its use  $\pm \omega$
- Motion of simple pendulum and relation.
- Kinetic energy and potential energy during Simple harmonic motion.
- Free, Forced and Damped Oscillations
- Resonance

## 6. WAVES

### COURSE CONTENT

- Progressive waves
- Transverse and longitudinal waves.
- Principle of superposition.
- Stationary waves and wavelength of sound waves in air columns and stretched strings
- Doppler's Effect
- Electromagnetic Spectrum.

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## 7. LIGHT

### COURSE CONTENT

- Interference of light waves, constructive and destructive interference.
- Young's Double Slit experiment, fringe spacing, dark and bright fringes.
- Diffraction (basic principle)
- Diffraction grating

## 8. HEAT & THERMODYNAMICS

### COURSE CONTENT

- Basic postulates of kinetic theory of gases
- Pressure exerted by a gas and derive the relation  $PV = \frac{Nm}{3} \langle v^2 \rangle$
- Equation of state for an ideal gas as  $PV = nRT$
- $PV = \frac{Nm}{3} \langle v^2 \rangle$  and  $PV = NkT$  and prove that  $K.E \propto T$  for a single molecule.
- Internal Energy.
- Specific Heat capacity.

## 9. ELECTROSTATICS

### COURSE CONTENT

- Coulomb's Law
- Electric field strength.
- $E = \frac{\Delta V}{\Delta d}$  to calculate the field strength.
- Electric field lines.
- $E = \frac{Q}{4\pi\epsilon_0 r^2}$  for the field strength
- Gravitational force and electric force.
- Electric potential
- Capacitance of a capacitor
- Energy stored in capacitor

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## 10. CURRENT ELECTRICITY

### COURSE CONTENT

- Current
- Ohm's Law
- Series and parallel Combination of resistors
- Resistance and resistivity
- Potential difference and e.m.f
- Power dissipation in resistors.
- Kirchhoff's First Law as conservation of charge.
- Kirchhoff's Second Law as conservation of energy.
- Potentiometer

## 11. ELECTROMAGNETISM

### COURSE CONTENT

- Magnetic field
- Force on current carrying conductor in uniform magnetic field
- Force on a moving charge in magnetic field
- Motion of charge particle in uniform electric and magnetic field.
- e/m for an electron.

## 12. ELECTROMAGNETIC INDUCTION

### COURSE CONTENT

- Magnetic flux
- Faraday's Law and Lenz's Law.
- Induced e.m.f and factors
- Alternating current and use  $V = V_0 \sin \omega t$ .
- Transformer and uses  $\frac{N_s}{N_p} = \frac{V_s}{V_p} = \frac{I_p}{I_s}$  and practical transformer.
- Period, frequency, peak value and root mean square value of an alternating current or

- 
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voltage.

### 13. DEFORMATION OF SOLIDS

#### COURSE CONTENT

- Stress, strain and Young's Modulus.
- Tensile stress and strain.
- Hook's Law.
- Elastic and plastic deformation of a material.
- Strain energy.
- Band Theory, valence band, conduction band and forbidden band.

### 14. ELECTRONICS

#### COURSE CONTENT

- Half and Full wave rectification
- Single diode for half wave rectification of an alternating current
- Four diodes for full wave rectification of an alternating current
- Operational amplifier and its characteristics

### 15. MODERN PHYSICS

#### COURSE CONTENT

- Energy of photon  $E = hf$
- Photoelectric Effect, Threshold Frequency and Work Function Energy.
- Maximum photoelectric energy is independent of intensity whereas photoelectric current is proportional to intensity
- Einstein's Photoelectric equation  $hf = \phi + \frac{1}{2} m v_{\max}^2$
- de Broglie wavelength and use  $\lambda = \frac{h}{p}$ .
- Discrete energy levels of hydrogen atom and spectral lines.
- Relation  $hf = (E_2 - E_1)$

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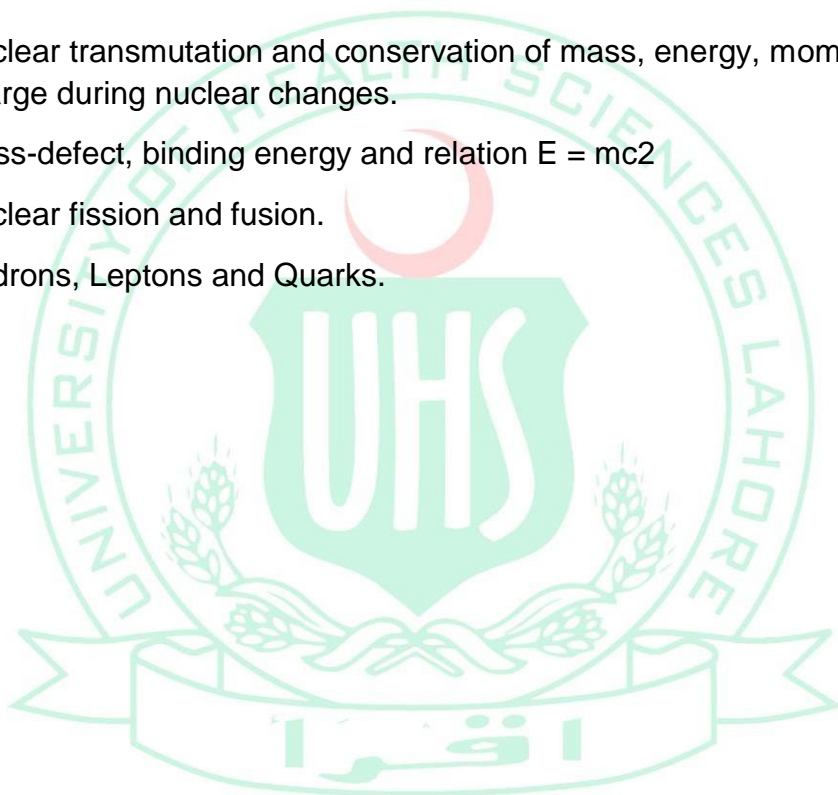
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- Production of X-rays and features of X-rays tube.

## 16. NUCLEAR PHYSICS

### COURSE CONTENT

- Nucleus, nucleon number and charge number.
- Radioactivity and emission of radiation.
- Activity, Decay constant and relation Activity =  $N\lambda$ .
- Half-life of radioactive substance and relation  $\lambda = \frac{0.693}{t_{1/2}}$
- Nuclear transmutation and conservation of mass, energy, momentum and charge during nuclear changes.
- Mass-defect, binding energy and relation  $E = mc^2$
- Nuclear fission and fusion.
- Hadrons, Leptons and Quarks.



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# ENGLISH

## Aim

The aim of the English section of MDCAT is to measure the applicants' skills in English language and to evaluate how prepared they are for undertaking graduate studies in medicine in English. The test applies a common standard to everyone to be able to evaluate the preparation of the applicants from different sectors, regions and socio-economic backgrounds.

The benchmarks for the test have been developed in the light of the curriculum used in HSSC

and CIE. Since the students who take the MDCAT come from a wide range of educational

contexts, the test comprises items that may be applied to a broadband of language

competencies that are not exclusive to one particular type of

## Objectives

- i. To ensure complete alignment between the English curriculum used in various sectors at the HSSC and CIE level and the test items
- ii. To create a balance of items from different benchmarks of the English curriculum outlined for MDCAT
- iii. To make sure that difficult and ambiguous items beyond the scope of high school education are not included
- iv. To design the test specifications
- v. To design, select, and arrange test task items

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Objective	Benchmark	Contents
1. Comprehend key vocabulary	Use one or more of the following strategies to determine meaning of key vocabulary: <ol style="list-style-type: none"> <li>1.1 contextual clues and illustrations</li> <li>1.2 background or prior knowledge</li> <li>1.3 morphology, syntax, phonics, knowledge of word relationships</li> <li>1.4 knowledge of synonyms, antonyms, homophones</li> </ol>	High and low frequency words from the course book or to be selected from similar contexts or the contexts the HSSC and CIE students may be familiar with
2. Demonstrate control of tenses and sentence structure	<ol style="list-style-type: none"> <li>2.1 Use correct tenses and sentence structure in writing</li> <li>2.2 Identify mistakes in the use of tenses and sentence structure in written texts</li> </ol>	<ul style="list-style-type: none"> <li>• All the present tenses</li> <li>• All the past tenses</li> <li>• Four types of sentences</li> <li>• Conditionals</li> <li>• Types of clauses</li> <li>• Fragments</li> </ul>
3. Demonstrate correct use of subject-verb agreement	<ol style="list-style-type: none"> <li>3.1 use correct subject-verb agreement in written texts</li> <li>3.2 Identify mistakes in the use of subject verb-agreement in written texts</li> </ol>	Use the texts prescribed/used in HSSC or CIE for selecting test items as well as determining the degree of their complexity
4. Demonstrate correct use of articles and prepositions	<ol style="list-style-type: none"> <li>4.1 Use appropriate articles and prepositions in different written contexts</li> <li>4.2 Identify mistakes in the use of articles and prepositions in sentences or short texts</li> <li>4.3 Select the appropriate article or preposition for a particular context</li> </ol>	The test items to be selected from the contexts common to the texts at HSSC and CIE level
5. Demonstrate correct use of writing conventions of spelling, capitalization and punctuation to clarify meaning	<ol style="list-style-type: none"> <li>5.1 Use capitalization and punctuation such as semi colons, commas in a series, apostrophes in possessives, proper nouns, and abbreviations</li> <li>5.2 Avoid and identify the following punctuation</li> </ol>	The test items to be selected from the type of texts written by HSSC and CIE students and from the contexts common to both the streams
	mistakes in sentences or short written texts: <ol style="list-style-type: none"> <li>5.2.1. Run on sentences</li> <li>5.2.2. Comma splices</li> <li>5.2.3. Fragments</li> <li>5.2.4. Faulty coordination</li> </ol>	

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## Vocabulary

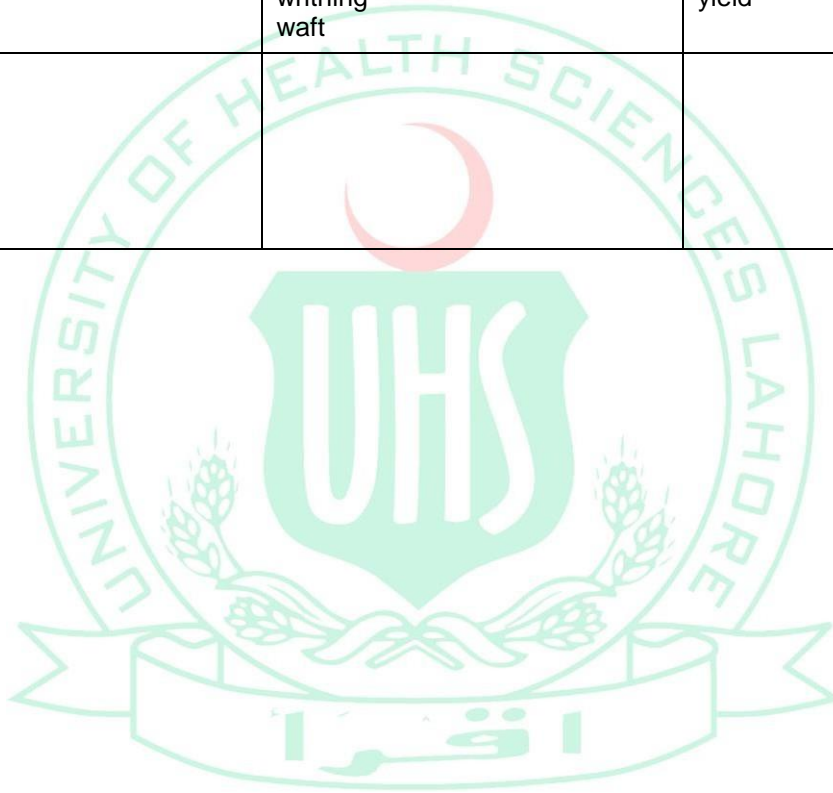
<p><b>A</b></p> <p>appalling astounded apparently attached to appraised alas abruptly accentuated anxieties</p>	<p><b>B</b></p> <p>boon bewilderment briskly bead brimmin g baffling bashful beckone d</p>	<p><b>C</b></p> <p>coveted credentials capacious collided with crudely confront compelled crudely coaxed comprehension curious casually confining crammed confirm cautioned captivated condescended compelled criteria</p>
<p><b>D</b></p> <p>daintily dispensing dispute distract drummed dilapidated disconsolately delicately dank dilapidated disguise definite</p>	<p><b>E</b></p> <p>enchanted encouraged fringed exude eccentric excursion elaborate exasperatio n expansive exaggeratio n evaluates</p>	<p><b>F</b></p> <p>fatality flicked flawlessly friction fluttered</p>
<p><b>G</b></p> <p>gingerly glistening glared groggy</p>	<p><b>H</b></p> <p>heap hideous habitat haggard haphazardl y harmony haughty havoc hearsay</p>	<p><b>I</b></p> <p>intended in vain illumination invariably irritable insinuated intently industry intolerable intolerable imperceptibly</p>
<p><b>J</b></p> <p>judgment judicial</p>	<p><b>K</b></p> <p>Kindred Knack</p>	<p><b>L</b></p> <p>likelihood labyrinth</p>

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<p>junction juvenile jeopardy jealousy jubilant</p>	<p>knickers knickknack</p>	<p>ludicrous limp</p>
<p><b>M</b></p> <p>menaced mustered mean mass mounting minimum mayhem miniature mumbled meditated moulded menacing</p>	<p><b>N</b></p> <p>nuisance naïve native nauseous negate negligence nemesis neutral niggle</p>	<p><b>O</b></p> <p>opted for on the wrong foot occasionally operation</p>
<p><b>P</b></p> <p>plopped presume precautions panting purchase persisted pensively prime placidly peered propelling passion promptly practically prone to paraphernalia prerogative path precision pizzazz potential</p>	<p><b>Q</b></p> <p>qualitative qualm quantitative quarrel quench query queue quirk quiver quizzical quotation</p>	<p><b>R</b></p> <p>rituals reinforce reprimanded riot reluctantly refuge regret rarely reproachful ragged revolving resonant</p>
<p><b>S</b></p> <p>swarmed up scenario swathe subsequently struck up string sternly</p>	<p><b>T</b></p> <p>tentatively tackle tumultuous tomfoolery tangle troughs tangled</p>	<p><b>U</b></p> <p>urge unburdened unprovoked</p>

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solemnly succulent shuffled sailed stunt sauntere d splendou r sagged off speckled with stable		
<b>V</b> ventured vulnerable	<b>W</b> whipped weighing up writhing waft	<b>Y</b> Yearnin g Yelp yield
<b>Z</b> zealou s zenith zest		



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